

CHEMISTRY WORKSHEET - 2  
CLASS - X

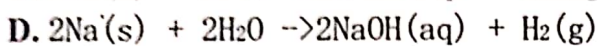
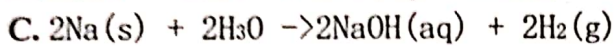
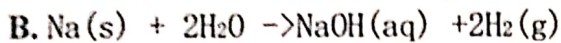
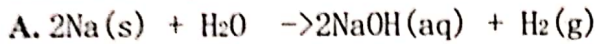
**Student Worksheet**

**Time: 30 minutes**

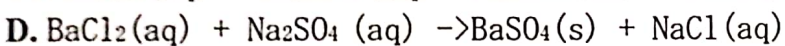
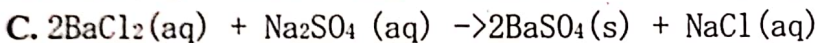
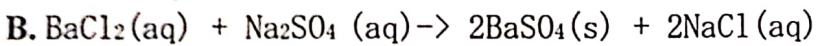
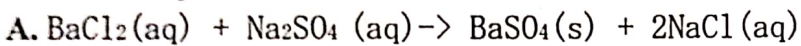
There are 15 multiple choice questions, only one of the options in every question is correct. Choose the correct option.

**Q1.** The given equation represents the reaction of sodium metal with water.  
Sodium + Water Sodium hydroxide + Hydrogen.

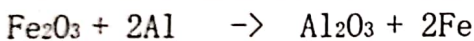
Which of the following chemical equations represents a complete balanced chemical equation for the given word equation?



**Q2.** Identify the chemical equation which represents a complete balanced equation for the reaction of barium chloride with sodium sulphate to produce barium sulphate and sodium chloride.



**Q3.** Ferric oxide reacts with aluminum to produce aluminum oxide and iron. The balanced chemical equation for the given reaction is

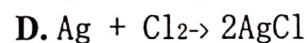
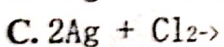
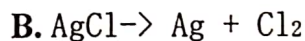
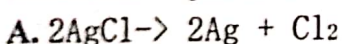


Ferric oxide Aluminum Aluminum oxide Iron

Which of the following substances is oxidized in the given reaction?

- A.  $\text{Al}_2\text{O}_3$     B.  $\text{Fe}_2\text{O}_3$     C. Al    D. Fe

**Q4.** White silver chloride changes to grey in sunlight because of the formation of silver metal along with the evolution of chlorine gas. Identify the chemical equation that gives a correct representation of the process.

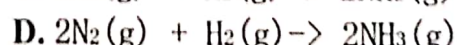
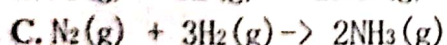
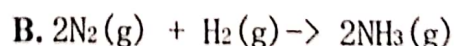
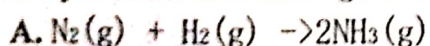


**Q5.** When copper oxide is treated with hydrogen gas copper is produced along with water. The balanced chemical equation for the given reaction is

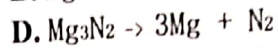
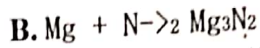
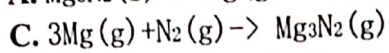
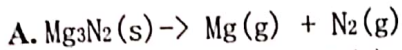
$\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$ . Which substance is oxidized in the given reaction?

- A.  $\text{H}_2$     B. Cu    C.  $\text{H}_2\text{O}$     D. CuO

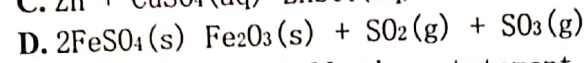
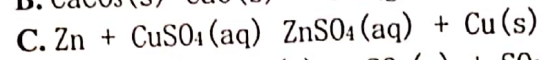
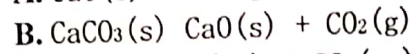
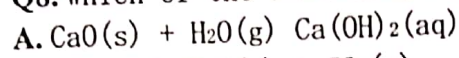
**Q6.** The balanced chemical equation for the reaction of nitrogen with hydrogen to produce Ammonia gas is-



**Q7.** Identify the chemical equation that correctly represents production of magnesium nitride by burning magnesium metal in a pure nitrogen atmosphere from the following-



Q8. Which of the following reactions represents a combination reaction?



Q9. Complete the following statement by choosing correct type of reaction for x and y.

Statement 1: The heating of lead nitrate is an example of 'x' reaction

Statement 2: The burning of magnesium is an example of 'y' reaction."

A. x y.

combination decomposition

B. x y

decomposition combination

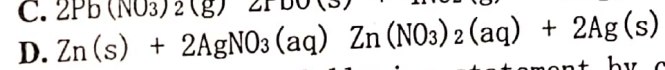
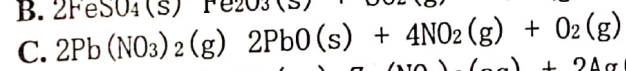
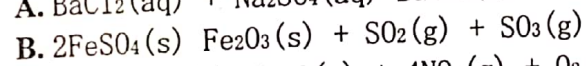
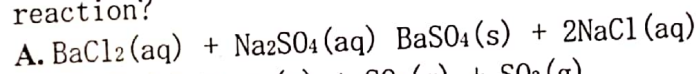
C. x y

combination displacement

D. x y

displacement decomposition

Q10. Which of the following reactions represents a double displacement reaction?



Q11. Complete the following statement by choosing correct options for x and y.

"During the process of respiration, glucose combines with oxygen in the cells of our body and 'x' a large amount of energy. Hence, respiration is an 'y' process."

A. x y

releases endothermic

B. x y

absorbes endothermic

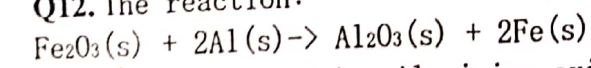
C. x y

releases exothermic

D. x y

absorbes exothermic

Q12. The reaction:



Ferric oxide Aluminium Aluminium oxide Iron

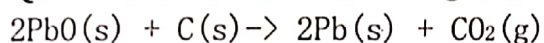
is an example of a-

- A. combination reaction B. double displacement reaction  
C. decomposition reaction D. displacement reaction

Q13. Complete the following statement by substituting x and y with correct options "Corrosion and rancidity are the result of 'x' reaction of iron articles and oils/fats respectively. Galvanization is done to prevent corrosion of iron articles and 'y' are used to prevent rancidity of oils or fats."

- A. x y  
displacement oxidants  
B. x y  
oxidation anti-oxidants  
C. x y.  
displacement anti-oxidants  
D. x y  
oxidation anti-oxidants

Q14. Which of the following statements about the reaction below is incorrect?



- (i) Lead is getting reduced (ii) Carbon dioxide is getting oxidized  
(iii) Carbon is getting oxidized (iv) Lead oxide is getting reduced  
A. (i) and (ii) B. (iii) and (iv)  
C. (i), (ii) and (iii) D. all

Q15. What happens when dilute hydrochloric acid is added to iron fillings?

- A. Hydrogen gas and iron chloride are produced  
B. Chlorine gas and hydroxide are produced  
C. heat is absorbed, i.e. test tube becomes cold.  
D. Iron salt and water are produced